Code No: 152AB

R18

JAWAHARLAL NEHRU TECHNOLOGICAL UNIVERSITY HYDERABAD B. Tech I Year II Semester Examinations August

CHEMISTRY	
(Common to CE, ME, ECE, EIE, MCT, MMT, AE, MIE, PTM)	
Note: This question paper contains two parts A and B. Part A is compulsory which carries 25 marks. Answer all questions in Part B consists of 5 Units. Answer any one full question from each unit.	art A
question carries 10 marks and may have a, b, c as sub questions.	Laci
, was any of a safe discipline.	
h) Write the molecular orbital electronic configurations of O ₂ . How can we recover exhausted ion exchange resins? What is battery? And discuss the types of batteries. Define the terms enantiomers and diastereomers. What are the conditions for molecule to be IR active? Define conductors and semi conductors on the basis of Band theory. A sample of hard water contains the following dissolved salts per fitre. Ca (HCO ₃) ₂ = 16.4 mgs, Mg(HCO ₃) ₂ = 14.6 mgs, CaCl ₂ = 111 mgs, MgSO ₄ = 12 mgs, CO ₂ = 44 mgs and CaSO ₄ = 13.6 mgs. Calculate temporary and permanent hardness of water in ppm and in degree Clarke. b) Explain how nature of metals effects the rate of corrosion What are starged and a last in the following dissolved salts per fitre. [3]	s] i]
j) What do you mean by absorbance and transmittance? [3	T 1 10
-with with war income one of	1
PART-B	
(50 M)	arks)
 Explain the crystal field theory for the square planar complexes of transition metals. Draw the molecular orbital energy diagram of F₂ molecule. Calculate its Bond orde predict its magnetic behavior. 	r and
	3+3]
3.a) What is doping agent? Explain how the dopant will effect the conductivity of conductors.	semi
b) Discuss the reasons for crystal field splitting of d-orbitals. Explain the splittin d-orbitals in octahedral complexes.	g of
c) Discuss the assumption of molecular orbital diagram. [4+3	3+3]
.a) Discuss the principle involved in EDTA method? Explain Estimation of hardnes water by complex metric method.	s of

- 0.28 g of CaCO3 was dissolved in dil.HCl and the solution made up to one litre with distilled water. 100 ml of the above solution required 28 ml of EDTA solution for titration. 100ml of the water sample required 33 ml of same EDTA solution for titration. After boiling, 100ml of this water, cooling, filtering and then titration required 10ml of EDTA solution. Calculate the temporary and permanent hardness in water.

c) Write a brief account on break-point chlorination.

[4+3+3]

- 5.a) What is the principle involved in internal treatment of boiler feed water and discuss the types of internal treatment methods.
 - b) What is hardness of water due to? What are its units? Explain the different types of hardness of water?
 - c) Explain the desalination of water by Reverse osmosis process.
- 6.a) What is a reference electrode? Explain the functioning of Quinhydrone electrode with Nernst equation.
 - b) What is electrochemical series? Write its applications.
- c) Give a brief account on lead –acid storage cell. [4+3+3]

OR

- 7.a) What is oxidation corrosion and how does it take place? Describe the mechanism of oxidation corrosion.
 - b) Explain impressed current method of corrosion control with a neat diagram.
- c) Write a note on galvanizing. [4+3+3]
- 8.a) Explain different conformations of n-Butane with the potential energy diagram.
- b) Give the structure, synthesis and applications of paracetamol.
- c) Why does SN^2 reactions give rise to inversed product? Explain. [4+3+3]

OR

- 9.a) Give Grignard addition reactions on carbonyl compounds.
 - b) What is markownikoff rule? Explain with suitable- example.
 - c) Discuss the mechanism involved in oxidation of alcohols using $KMnO_4$. [4+3+3]

[4+3+3]

- 10.a) Discuss the type electronic transitions in U.V spectra?
 - b) Calculate the theoretical number of vibrational degree of freedom of
 - i) CO₂ ii) Water iii) SO₂
 - c) Predict the NMR spectrum of

i) i

ii)

 H_2C H_2 CH_3

OR

- 11.a) What is meant by term chemical shift. Explain the factors influencing chemical shift.
 - b) Discuss the applications of UV Spectroscopy.
 - c) Give an account of various fundamental vibrations in IR spectra. [4+3+3]